Sp	pecific Heat Capacity N	ame
$E = m c \theta$		
Ε	Energy (in/out) = mass x specific heat cap	acity x temperature change
Specific heat capacity of water is 4200 J/kg°C; aluminium is 900 J/kg°C; glass is 500 J/kg°C		
1.	How much energy is needed to heat 2kg of water from	
2.	Calculate the energy needed to raise the temperature of a) 100g of water by 5°C	f
	b) 500g of water by 4°C	
	c) 800g of aluminium by 20°C	

A 1.5 kg piece of iron absorbs 100 kJ of heat energy, and its temperature changes from 25°C to 175°C. Calculate the specific heat capacity of iron.

3.

4.	How many joules of heat are needed to raise the temperature of 0.1 kg of aluminum from 22°C to 55°C?
5.	A 50 g piece of glass is at 20°C. What will its temperature rise to if it absorbs 5275 joules of heat?
6.	Calculate the heat capacity of a piece of wood if 3 kg of the wood absorbs 15kJ of heat, and its temperature changes from 32°C to 57°C.
7.	100mL of water at 4°C is heated until its temperature is 37°C. Calculate the amount of heat energy needed to cause this rise in temperature.